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A compound analyzed in a chemistry lab consists of 5.34 g of carbon, 0.42 g of Hydrogen, and 47.08 of

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Chlorine.

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STUDY. PLAY. Mole.

(mol) of a substance is
6.02 times 10^{23}

representative

particles of that

substance and is the SI

unit for measuring the

amount of a substance.

Avogadro's Number.

the number of

representative

particles in a mole,

6.02 times 10^{23} .

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Representative
Particle.

Quantities
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STUDY. PLAY. mole.

6.02×10^{23}

representative

particles of a

substance, SI unit for

measuring the amount

of a substance.

Avogadro's number.

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6.02×10^{23}

representative

particles of a

substance, named in

honor of Italian

scientist Amedeo

Avogadro di Quaregna.

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Chapter 10. Avogadro's
number. Empirical
Formula. Molar Mass.
Molar Volume. number
of particles in one mole
of a pure substance

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Chapter 10
(element o.... Formula
that shows the lowest
whole-number ratio of
the atoms.... The mass
of one mole of an
element. Found on the
periodic tabl....

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more with flashcards,
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...

10.2 Mole to Mole &
Mole to Volume
Relationships *For all
the problems on this
page, first find the
molar mass of the
compound. 1. What is
the mass of 9.45 mol of
aluminum oxide? mol A

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Chapter 10

ILO + 3 (I A GO q

63,5ù?- 2. What is the mass of 4.52×10^{-3} mol of ethylbenzene, $C_6H_5CH_2CH_3$?

*Ethylbenzene is a hydrocarbon that is produced by burning coal.

BHS - Moodle

A chemistry student working in the lab might be asked to calculate how much 1-bromo-2-methylpropane, C_4H_9Br , could be

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Chapter 10
made from 6.034 g of
... 10 1 kg 103 g 368
Chapter 10 Chemical
Quantities
Calculations and
Chemical Equations.
10.1 Equation
Stoichiometry 369 The
ratio of moles of P 40

Chapter 10 Chemical Calculations and equations

6.7 Chapter Summary.
To ensure that you
understand the
material in this
chapter, you should

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review the meanings of the following bold terms in the following summary and ask yourself how they relate to the topics in the chapter. Chemical reactions relate quantities of reactants and products.

Chapter 6 - Quantities in Chemical Reactions - Chemistry

Chapter 10 - Chemical
Quantities, 10.1 The

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Chapter 10
Mole: A Measurement
of Matter - Sample
Problem 10.1. 10.1 The
Mole: A Measurement
of Matter - Chemistry &
You. 10.1 The Mole: A
Measurement of Matter
- Sample Problem 10.2.
10.1 The Mole: A
Measurement of Matter
- Sample Problem 10.3.

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Mass and Mole-Volume

Relationships - 10.2

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SI unit representing

6.02×10^{23}

representative

particles of a

substance. the

temperature and

pressure at which one

mole of gas occupies a

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volume of 22.4 L. equal volumes of gases at the same temperature and pressure contain equal numbers of particles.

Quia - Chapter 10 "Chemical Quantities" Vocab

Use the chemical formula to find the number of atoms in one molecule and multiply this number by Avogadro's number, the number of particles

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in one mole. atom 6.02

• 10²³ O 2 6.02 • 10

23 ion Na⁺ 6.02 •

10²³ formula unit NaCl

6.02 • 10²³ 6.02 •

10²³ representative

particles of a

substance molecule

formula unit atom

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- Mr. Mutic's

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Quantities
Get Free Chapter 10
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Guided Practice
Answers substance is
called a mole • A
mole(mol) of a
substance is equivalent
to 6.02×10^{23}
particles of that
substance • The mole
was founded by a
scientist named
Avagadro, and he
decided to use the

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Chapter 10
CHAPTER 10: Chemical
Quantities Chapter 10
Chemical Quantities 91
SECTION 10.1 THE ...

Chapter 10 Chemical Quantities Guided Practice Answers

Here is an example.
Pure HCl (hydrogen
chloride) is a gas that
is very soluble in
water. A plot of the
partial pressure of
gaseous HCl in
equilibrium with
aqueous HCl, as a

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function of the solution molality (Fig. 10.1), shows that the limiting slope at infinite dilution is not finite, but zero.

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