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Calculate Kc at this temperature. CH3COOH(l) + CH3CH20H(l) CH3COOCH2CH3(l) + H20(l) .1.00 a-. 00. ..v.s.s.s o. XY<S. +0. (4) (A2 BBoF 29) 1.90 mol of hydrogen and 1.90 mol of iodine were allowed to reach equilibrium at 710 K. The equilibrium mixture contained 3.00 mol of HI.

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Chemsheets AS 1044 (Kc Calculations 2) Chemsheets AS 1043 (Kc Calculations 1) Chemsheets AS 1042 (Equilibrium quantities 2) Chemsheets AS 1041 (Equilibrium quantities 1) Chemsheets AS 1040 (The Haber Process) Chemsheets AS 1039 (Le Chatelier 2) Chemsheets AS 1038 (Le Chatelier 1) 3.1.6 Chemical equilibria, Le Chatelier_s principle BOARDWORKS ANSWERS BELOW Chemsheets AS 1044 (Kc…

3.1.6 Chemical equilibria, Le Chatelier’s principle and Kc ...
There are all sorts of calculations you might be expected to do which are centred around equilibrium constants. You might be expected to calculate a value for K c including its units (which vary from case to case). Alternatively you might have to calculate equilibrium concentrations from a given value of K c and given starting concentrations.

equilibrium constants - Kc
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1) CH 4(g) + H 2 O (g) CO (g) + 3 H 2(g) initial moles 1 1 0 0 change in moles equilibrium moles 0.4 9) 2) N 2(g) + 3 H 2(g) 2 NH 3(g) initial moles 0.1 0.1 0

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initial moles 1 1 1 1 change in moles -0.2 -0.2 +0.2 +0.6 equilibrium moles 0.8 0.8 1.2 1.6 8) N 2(g) + 3 H 2(g) 2 NH 3(g) initial moles 10 0 10 change in moles +0.5 +1.5 -1,0

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change in moles -1.4 -1.4 +2.8 equilibrium moles 0.6 0.6 2.8 d) H 2(g) + I 2(g) ⇌ 2 HI (g) initial moles 1 5 1 change in moles -0.2 -0.2 +0.4 equilibrium moles 0.8 4.8 1.4 e) 2 SO 2(g) + O 2(g) ⇌ 2 SO 3(g)

EQUILIBRIUM QUANTITIES 2
Le Chatelier 1 (1038) WS: Join: Join: Le Chatelier 2 (1039) WS: Join: Join: The Haber process (1040) WS: Join: Join: Equilibrium constant Kc (1258) PP: Join: Equilibrium quantities 1 (1041) WS: Join: Join: Equilibrium quantities 2 (1042) WS: Join: Join: Kc and its units (1174) WS: Join: Join: Kc calculations 1 (1043) WS: Join: Join: Kc ...

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reacting mass calculations 1 1) Aluminium is extracted from aluminium oxide as shown. Calculate the mass of aluminium that can be formed from 1020 g of aluminium oxide.

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In combustion studies of H2 as an alternative fuel, you find evidence that the hydroxyl radical (HO) is formed in flames by the following reaction. H(g) + 1/2 O2(g) <---> HO(g) Use the following data to calculate Kc for this reaction. 1/2 H2(g) + 1/2 O2(g) <---> HO(g) Kc = 0.58 1/2 H2(g) <---> H(g) Kc = 1.6 X 10-3

Kc Calculation? | Yahoo Answers
1) Calculate the pH of the following solutions. a) 0.15 mol dm-3 KOH b) 0.05 mol dm-3 NaOH c) 0.20 mol dm-3 Ba(OH) 2 2) Calculate the concentration of the following acids. a) NaOH with pH 14.30 b) Ba(OH) 2 with pH 12.50 c) KOH with pH 13.70 3) Calculate the pH of the solutions formed in the following way.

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Ice Table - Equilibrium Constant Expression, Initial Concentration, Kp, Kc, Chemistry Examples - Duration: 39:00. The Organic Chemistry Tutor 672,600 views 39:00