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The kinetic theory of gases is a historically significant, but simple, model of the thermodynamic behavior of gases, with which many principal concepts of thermodynamics were established. The model describes a gas as a large number of identical submicroscopic particles (atoms or molecules), all of which are in constant, rapid, random motion. Their size is assumed to be much smaller than the ...

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Kinetic theory of gases - Wikipedia

1. kinetic theory Of matter 2. plasma 3. crystals 4. solid rżóamorphous material 6. steam _e_ 7. thermal expansion 8. liquid 10. ice use the words in the box to fill in the blanks. Column II water vapor b. state of matter with no definite shape but with definite volume c. solid which is not made of crystals d.

Quia

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Section 1 Reinforcement Earth Answers

States of Matter Section 13.1 Gases In your textbook, read about the kinetic-molecular theory. Complete each statement. 1. The kinetic molecular theory describes the behavior of gases in terms of particles in . 2. The kinetic-molecular theory makes the following assumptions. a. In a sample of a gas, the volume of the gas particles themselves is very

Study Guide for Content Mastery - Teacher Edition

Chapter 10: Section 1: The Kinetic-Molecular - Section 1: The Kinetic-Molecular Theory of Matter kinetic-molecular theory. Study Everywhere! 2015 Quizlet Inc. Follow @quizlet on Twitter. Presentation " Section 1. Kinetic Theory: How - Section 1. Kinetic Theory: How particles in matter behave 3 Basic Assumptions of the Kinetic Theory 1. All matter is composed of small particles (atoms,

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